

CHEMISTRY



Target : JEE (Main) &
Target : JEE (Main + Advanced)

BASIC INORGANIC NOMENCLATURE
WWW.GRADESETTER.COM

Contents

Topic	Page No.
Theory	01 - 06
Exercise - 1	07 - 08
Exercise - 2	08 - 10
Answer Key	11 - 12
Advanced Level Problems	13 - 14
Answer Key	15

Copyright reserved.

All rights reserved. Any photocopying, publishing or reproduction of full or any part of this study material is strictly prohibited. This material belongs to a enrolled student of RESONANCE only any sale/resale of this material is punishable under law, subject to Kota Jurisdiction only.



Th-2 Elements

General Rule : The names of metals generally end with-ium or-um (examples are sodium, potassium, aluminum, and magnesium).

The exceptions are metals that were used and named in ancient times, such as iron, copper, and gold.

The names of nonmetals frequently end with-ine, -on, or -gen (such as iodine, argon, and oxygen.)

Given the names of the constituent elements and common ions, most of the common inorganic compounds can be named using the rules presented below.

Th-3 Acids :

Acids are normally classified in two groups, hydracids and oxyacids

Hydracids :

Hydracids are acids which contain hydrogen and a non-metal, but no oxygen.

General Rule : The names of hydracids have the prefix hydro-(sometimes shortened to hydr-) and the suffix-ic attached to the stem based on the names of the constituent elements (other than hydrogen.)

For example, HCl (made of hydrogen and chlorine) is hydrochloric acid; HBr (made of hydrogen and bromine) is hydrobromic acid; HI (made of hydrogen and iodine) is hydroiodic acid; HCN (made of hydrogen, carbon, and nitrogen) is hydrocyanic acid; and H_2S (made of hydrogen and sulfur) is hydrosulfuric acid.

Th-4 Oxoacids or Oxyacids

The acids which contain hydrogen, oxygen and a metal or non-metal.

In this case, more than one possibility arises due to the presence of different number of oxygen atoms. An example of such an oxoacid series is as follows: $HClO$, $HClO_2$, $HClO_3$, $HClO_4$. All these contain same three elements but differ in the number of oxygen atoms present.

General Rule-1 :

If a class of acids contains only one member, its name is given the suffix-ic.

For example, hydrogen, carbon and oxygen combine to form only one acid i.e. H_2CO_3 . It is called carbonic acid (carbonic acid.)

General Rule-2 :

If an acid series contains two acids, such as H_2SO_4 and H_2SO_3 , the acid containing more oxygen atoms is given the suffix -ic, while the acid with fewer oxygen atoms is given the suffix-ous.

For example, H_2SO_4 is sulphuric acid, and H_2SO_3 is sulphurous acid.

Similarly, HNO_3 is nitric acid and HNO_2 is nitrous acid.

General Rule-3 :

The prefix ortho and meta have been used to distinguish acids differing in the 'content of water'

(H_3BO_3) - orthoboric acid $-H_2O$

$(HBO_2)_n$ -metaboric acid

General Rule-4 :

The prefix pyro has been used to designate an acid formed from two molecules of an ortho acid minus one molecule of water.

ex- $H_4P_2O_7$ -pyro phosphoric acid

General Rule-5 :

The prefix peroxy indicates the substitution '-O-' by '-O-O-'

HNO_4 - peroxy nitric acid

H_3PO_5 -peroxy mono phosphoric acid

General Rule-6 :

Acid derived by oxoacids by replacement of oxygen by sulphur are called thio acids.

$H_2S_2O_2$ - thio sulphurous acid

$H_2S_2O_3$ - thio sulphuric acid

Note : when more than one oxygen atom can be replaced by sulphur the number of sulphur atom should generally indicated

H_3PO_3S mono thio phosphoric acid

$H_3PO_2S_2$ Dithiophosphoric acid

In the case of an extensive acid series (such as $HClO$, $HClO_2$, $HClO_3$, $HClO_4$), the acid with the one oxygen atoms lesser than -ous acid is given the prefix hypo- and the suffix -ous, and the acid with the one oxygen atom more than the -ic acid is given the prefix per and a suffix-ic.

In the above example, $HClO$ is hypochlorous acid, $HClO_2$ is chlorous acid, $HClO_3$ is chloric acid, and $HClO_4$ is perchloric acid.



Resonance
Educating for better tomorrow

Corporate Office: CG Tower, A-46 & 52, IPIA, Near City Mall, Jhalawar Road, Kota (Raj.)-324005
Website : www.resonance.ac.in | E-mail : contact@resonance.ac.in
Toll Free : 1800 200 2244 | 1800 258 5555 | CIN: U80302RJ2007PLC024029



Scanned with
CamScanner

BIN - 3

Rule : Names of positive ions end with -ium if the ion has only one oxidation state (Only one level of net charge). For example, the positive ion of sodium is Na^+ (sodium ion), and the positive ion of aluminium is Al^{3+} (aluminium ion).

Metal atoms with more than one possible charges

Rule : If the cation has variable valency (charge), charge is specified in roman numerals in round brackets immediately after the name of metal atom. For example, Sn^{2+} is written as tin (II) ion.

Alternately, the less positive ion ends with -ous, and the more positive ion ends with -ic. For instance, the two positive ions of copper are Cu^+ (cuprous) and Cu^{2+} (cupric). The oxidation state of a positive ion can also be designated by placing a Roman numeral after the name of the elements. These positive ions of copper can also be written as copper(I) and copper(II), respectively.

Ions	Name
Cu^+	cuprous ion
Sn^{2+}	Stannous ion
Sn^{4+}	Stannic ion
Fe^{3+}	Ferric ion
Fe^{2+}	Ferrous ion

General Rule-3

Suffix -ium is often used with cations containing non metals.

For example, the positive ion of ammonia is NH_4^+ (ammonium) and the positive ion of water (H_2O) is H_3O^+ or H^+ (hydronium).

Remember these names !

NO_2^+ : nitronium

NO^+ : nitrosonium

H_3O^+ : hydronium

From NH_3 ammonia is derived NH_4^+ : ammonium.

Similarly.

N_2H_4 : hydrazine \rightarrow N_2H_5^+ : hydrazinium

$\text{C}_6\text{H}_5\text{NH}_2$: aniline \rightarrow $\text{C}_6\text{H}_5\text{NH}_3^+$: anilinium

$\text{C}_5\text{H}_5\text{N}$: pyridine \rightarrow $\text{C}_5\text{H}_5\text{NH}^+$: pyridinium

Th-6 Anions (Negative Ions)

Anions can always be looked upon as ions derived from acids by removal of one or more protons. Accordingly, anions can be classified as follows :

Anions derived from hydracids

Rule : Names of negative ions from hydracids end in -ide.

For example, Cl^- (chloride) from HCl , and CN^- (cyanide) from HCN . Following examples will give you a better insight in this nomenclature. It is also useful to remember them.

Remember these names

Anion	Name
H^-	Hydride ion
D^-	Deuteride ion
F^-	Fluoride ion
Cl^-	Chloride ion
Br^-	Bromide ion
I^-	Iodide ion
O^{2-}	Oxide ion
S^{2-}	Sulphide ion
Se^{2-}	Selenide ion
Te^{2-}	Telluride ion
N^{3-}	Nitride ion
P^{3-}	Phosphide ion
As^{3-}	Arsenide ion
Sb^{3-}	Antimonide ion
C^{4-}	Carbide ion
Si^{4-}	Silicide ion
B^{3-}	Boride ion

Anions derived from oxyacids

Anion derived from an oxyacid by removal of one or more H^+ ions is termed as oxyanion.

Rule : If the oxyacid is -ic acid, suffix -ate is used with oxy-anion.

For example

CO_3^{2-}	carbonate (from H_2CO_3)
ZnO_2^{2-}	zincate
SiO_3^{2-}	silicate

Rule : If the oxyacid is -ous acid, suffix -ite is used with oxy-anion.

For example

NO_2^- (nitrite) is derived from HNO_2 (nitrous acid), and SO_3^{2-} (sulphite) is derived from H_2SO_3 (sulphurous acid)

Rule : If the oxyacid has prefixes per- or hypo-, the oxyanion will have same prefixes.

For example, ClO_4^- perchlorate ion from HClO_4 , perchloric acid, ClO^- hypochlorite ion from HClO , hypochlorous acid

Remember these names !

SO_4^{2-}	sulphate
SO_3^{2-}	sulphite
NO_3^-	nitrate,
NO_2^-	nitrite
SnO_3^{2-}	stannate
SnO_2^{2-}	stannite,
PbO_3^{2-}	plumbate,
PbO_2^{2-}	plumbite

Anions containing replacable hydrogen ions

Polyprotic acid. Any acid containing more than one replacable hydrogens is said to be a polyprotic acid.



Basic Inorganic Nomenclature

Replacable hydrogens. H atoms which can be lost as H⁺ in reactions with a base.

H atoms connected to O atoms in oxyacids are all replacable. If all the replacable hydrogens are removed, we obtain the anions discussed in the sections above.

However, in all the polyprotic acids it is always possible to remove less than the maximum number of replacable hydrogens. e.g. H₃PO₄ is triprotic. We can remove one, two or three H⁺ ions from it to generate H₂PO₄⁻, HPO₄²⁻ and PO₄³⁻.

You are already familiar with phosphate ion, PO₄³⁻. The other two anions, H₂PO₄⁻ and HPO₄²⁻ still contain H atoms that are replacable. We consider their nomenclature in this section.

Rule-1: A prefix bi- (old notation) or hydrogen- (IUPAC notation) is attached to the name of anion.

Rule-2: For triprotic or higher acids, numerical prefixes (e.g. mono, bi, tri) are also used to indicate the number of replacable H atoms left in the sample.

eg. HCO₃⁻ is bicarbonate or hydrogen carbonate
 HSO₃⁻ bisulphite or hydrogen sulphite
 HS⁻ bisulphide or hydrogen sulphide etc.
 when anion has -3 charge, e.g. PO₄³⁻ then following possibilities arise.
 HPO₄²⁻ monohydrogen phosphate, H₂PO₄⁻ dihydrogen phosphate.

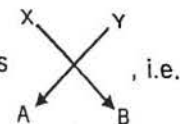
Miscellaneous Anions (To be comitted to memory)

Anion	Name
HO ⁻	Hydroxide ion
O ₂ ²⁻	Peroxide ion
O ₂ ⁻	Superoxide ion
S ₂ ²⁻	Disulphide ion
I ₃ ⁻	Triiodide ion
N ₃ ⁻	Azide ion
NH ²⁻	Imide ion
NH ₂ ⁻	Amide ion
CN ⁻	Cyanide ion
C ₂ ²⁻	acetylide ion
O ₃ ⁻	Ozonide ion
MnO ₄ ²⁻	Manganate ion
MnO ₄ ⁻	Permanganate ion
SCN ⁻	Thiocyanate ion
S ₂ O ₃ ²⁻	Thiosulphate ion
H ₃ COO ⁻	Acetate ion
C ₂ O ₄ ²⁻	Oxalate ion

Th-7 Method of writing formula of an ionic compound

In order to write the formula of an ionic compound which is made up of two ions (simple or polyatomic) having net charges x and y respectively, follow the following procedure.

- (i) Write the symbols of the ions side by side in such a way that positive ion is at the left and negative ion at the right as AB.
- (ii) write their charges on the top of each symbol as A^xB^y.



(iii) Now apply criss- cross rule as

formula A_yB_x.

(iv) Cancel out any common factor (or HCF).

Examples :

1. Calcium chloride = CaCl₂
2. Aluminium oxide = Al₂O₃
3. Potassium phosphate = K₃PO₄
4. Magnesium nitride = Mg₃N₂
5. Calcium oxide = Ca₂O₂
6. Ammonium sulphate = (NH₄)₂SO₄

Cancelling the common factor, answer is CaO

Table-2 : Difference between Atoms and ions

	Atoms		Ions
1	Atoms are perfectly neutral	1	Ions are charged particles containing one or more atoms.
2	In atoms, the number of protons is equal to the number of electrons. Na (protons 11, electrons 11); Cl (protons 17, electrons 17)	2	In cations (positively charged ions), number of protons is more than the number of electrons. In anions (negatively charged ions) the no. of protons is less than the number of electrons. e.g. Na ⁺ (protons 11, electrons 10). Cl ⁻ (protons 17, electrons 18)
3	Except noble gases, atoms have less than 8 electrons in the outermost orbit e.g. Na : 2, 8, 1; Ca : 2, 8, 8, 2 Cl : 2, 8, 7; S : 2, 8, 6	3	Ions have generally 8 electrons in the outermost orbit, i.e., ns ² np ⁶ configuration. Na ⁺ : 2, 8; Cl ⁻ : 2, 8, 8; Ca ²⁺ : 2, 8, 8
4	Chemical activity is due to loss or gain or sharing of electrons as to acquire noble gas configuration	4	The chemical activity is due to the charge on the ion. Oppositely charged ions are held together by electrostatic forces.

Table - 3 : Naming of Oxyacid

Acid end with IC suffix	Suffix-ous		Prefix -per ; suffix-ic		Name
	Formula	Name	Formula	Name	
H ₃ BO ₃	HNO ₂	Nitrous Acid	HNO ₄	Peroxynitric acid	Pyrophosphoric acid
H ₂ CO ₃	H ₂ SO ₃	Sulphurous acid	H ₃ PO ₅	Peroxy monophosphoric acid	Pyrophosphorous acid
HONC	H ₂ S ₂ O ₅	Disulphurous acid	H ₄ P ₂ O ₅	Peroxy diphosphoric acid	Pyroboric acid
HOCN	HClO ₂	Chlorous acid	H ₂ SO ₅	Peroxy monosulphuric acid	Pyro silicic acid
HNO ₃	Prefix - Hypo ; suffix -ic		H ₂ S ₂ O ₈	Peroxy disulphuric acid	Pyrosulphuric acid
H ₂ NO ₂	Hypo nitrous acid		HClO ₄	Perchloric acid	
H ₃ PO ₄	Hypo chlorous acid		Prefix -thio		
H ₂ SO ₄	Hypo sulphurous acid		H ₂ S ₂ O ₃	Thio sulphuric acid	
HClO ₃	Prefix -meta ; suffix -ic		H ₂ S ₂ O ₂	Thio sulphurous acid	
H ₂ S ₂ O ₆	(HBO ₂) _n		H ₂ S ₂ O ₆	Dithionic acid	
	(HPO ₃) _n		H ₂ S ₂ O ₄	Dithionous acid	



Exercise-1

Marked Questions may have for Revision Questions.

PART - I : SUBJECTIVE QUESTIONS

- (a) Group the following elements (atomic numbers given) into various blocks noted below:
12, 19, 17, 25, 31, 42, 54, 23, 38

(i) s-blocks (ii) p-blocks (iii) d-blocks

(b) Which of the following are transition elements?
K, Mn, Ca, Cs, Fe, Cu, Pb
- Write the group number & period number of following atomic numbers.
7, 15, 35, 55, 80, 70, 75, 90, 105, 18
- Which of the following elements have + 3 as most popular oxidation state?
F, Sc, Al, As, Ga, Cu, Zn, Cr, V, Cl, Xe
- (a) Write the name of following cations .
 NO_2^+ , NO^+ , H_3O^+ , NH_4^+ , N_2H_5^+ , $\text{C}_6\text{H}_5\text{NH}_3^+$, $\text{C}_5\text{H}_5\text{NH}^+$

(b) Write the name of following anions .
 F^- Cl^- Br^- Γ^- O^{2-} S^{2-} N^{3-} P^{3-} As^{3-} Cu^- H^- Au^-
 CO_3^{2-} ZnO_2^{2-} SiO_3^{2-} NO_2^- SO_3^{2-} ClO_4^- ClO^- SO_4^{2-} NO_3^- SnO_3^{2-} SnO_2^{2-} PbO_3^{2-} PbO_2^{2-}

PART - II : ONLY ONE OPTION CORRECT TYPE

- An element has atomic number is 29 . It belongs to -
(A) 4th period, group 11 (B) 5th period, group 10 (C) 4th period, group II B (D) 5th period, IB group
- Which of the following contains atomic number of only s-block
(A) 55,12,18,53 (B) 13,33,54,83 (C) 3, 20,55,87 (D) 22,33,55,66
- What is the atomic number of last member of the seventh period of the extended form of periodic table?
(A) 116 (B) 118 (C) 120 (D) 122
- The oxidation number and covalency of sulphur in the sulphur molecule (S_8) are respectively :
(A) 0 and 2 (B) + 6 and 8 (C) 0 and 8 (D) +6 and 2
- The oxidation number that iron does not exhibit in its common compounds or in its elemental state is -
(A) 0 (B) +1 (C) +2 (D) +3
- The oxidation state of nitrogen varies from :
(A) -3 to + 5 (B) 0 to +5 (C) -3 to 1 (D) +3 to +5
- When H_2SO_3 is converted into H_2SO_4 the change in the oxidation state of sulphur is from :
(A) 0 to + 2 (B) +2 to + 4 (C) +4 to +2 (D) +4 to + 6
- The halogen that shows same oxidation state in all its compounds with other elements is :
(A) I_2 (B) F_2 (C) Cl_2 (D) Br_2
- Most stable oxidation state of gold is -
(A) + 1 (B) +3 (C) +2 (D) zero
- The most stable oxidation state of chromium is -
(A) +5 (B) +3 (C) +2 (D) +4



11. Which can have both +ve and -ve oxidation states?
(A) F (B) I (C) Na (D) He
12. Which metal exhibits more than one oxidation states?
(A) Na (B) Mg (C) Al (D) Fe
13. The atomic number of an element which can not show the oxidation state of +3 is-
(A) 13 (B) 32 (C) 33 (D) 17
14. The most common oxidation state of an element is -2. The number of electrons present in its outer most shell is -
(A) 2 (B) 4 (C) 6 (D) 8
15. Conversion of $PbSO_4$ to PbS the oxidation number of sulphur in PbS is-
(A) -2 (B) +6 (C) +4 (D) -1
16. Oxidation state of oxygen in H_2O_2 is :
(A) -2 (B) -1 (C) +1 (D) +2
17. The oxidation number of phosphorus in $Mg_2P_2O_7$ is :
(A) +5 (B) -5 (C) +6 (D) -7
18. In the conversion of Br_2 to BrO_3^- , the oxidation state of bromine changes from-
(A) 0 to +5 (B) -1 to +5 (C) 0 to -3 (D) +2 to +5
19. Oxidation number of S in S_2Cl_2 is :
(A) +1 (B) +6 (C) 0 (D) -1
20. Which of the following element shows only -1 oxidation number in combined state :
(A) F (B) Cl (C) Br (D) I

PART - III : MATCH THE COLUMN

1. Column-I
- (A) Sulphurous acid
 - (B) Per oxo disulphuric acid
 - (C) Pyro sulphuric acid
 - (D) Peroxo mono sulphuric acid

- Column-II
- (p) $H_2S_2O_8$
 - (q) $H_2S_2O_7$
 - (r) H_2SO_3
 - (s) Sulphur O.S + 6

2. Match the column :
Column-I

- (A) HIO_2
- (B) $Mg(IO)_2$
- (C) HIO
- (D) $MgHPO_3$

- Column-II
- (p) Magnesium hydrogen phosphite
 - (q) Iodous Acid
 - (r) Magnesium hypoiodite
 - (s) Hypoiodous acid

Exercise-2

Marked Questions may have for Revision Questions.

PART - I : ONLY ONE OPTION CORRECT TYPE

1. Dichromate ion is :
(A) CrO_4^{2-} (B) $Cr_2O_7^{2-}$ (C) CrO_3 (D) Cr_2O_4
2. Dithionic acid in following compound is :
(A) $H_2S_2O_6$ (B) $H_2S_2O_4$ (C) H_2SO_5 (D) $H_2S_2O_3$

RESONANCE STUDY CENTRES (Self Owned)

KOTA (Head Office):

Pre-Engineering Division: JEE (Advanced)
Pre-Engineering Division: JEE (Main)
Pre-Medical Division: AIIMS/ AIPMT
Tel.: 0744-3012222, 3192222, 6635555
e-mail: contact@resonance.ac.in

Commerce & Law Program Division (CLPD)

Tel.: 0744-3192229, 6060663
e-mail: clpd@resonance.ac.in

PCCP/PSPD/MEs

Tel.: 0744-2434727, 8824078330, 3192223, 2440488
e-mail: pccp@resonance.ac.in

DLPD

Tel.: 0744-6635556, 3012222
e-mail: dlpd@resonance.ac.in

eLPD

Tel.: 0744-3058242
e-mail: elpd@resonance.ac.in

JAIPUR

Tel.: 0141-6060661/ 64, 3103666, 6060662/ 63
e-mail: jaipur@resonance.ac.in

BHOPAL

Tel.: 0755-3206353, 3192222, 3256353
e-mail: bhopal@resonance.ac.in

NEW DELHI

Tel.: 011-6060660/ 2/ 3/ 4/ 5/ 6/ 7
e-mail: delhi@resonance.ac.in

LUCKNOW

Tel.: 0522-3192222, 3192223/ 4, 6060660/ 61/ 62
e-mail: lko@resonance.ac.in

KOLKATA

Tel.: 033-3192222, 6060660/ 01/ 02
email: kolkata@resonance.ac.in

NAGPUR

Tel.: 0712-3017222, 3192222, 6060660
e-mail: nagpur@resonance.ac.in

NANDED

Tel.: 02462-250220, 606066
e-mail: nanded@resonance.ac.in

MUMBAI

Tel.: 022-31922222, 60606600
e-mail: andheri@resonance.ac.in

UDAIPUR

Tel.: 0294-6060660, 5107510, 3192222
e-mail: udaipur@resonance.ac.in

BHUBANESWAR

Tel.: 0674-3192222, 3274919, 6060660/ 61
e-mail: bbsr@resonance.ac.in

AHMEDABAD

Tel.: 079-31922222/ 3/ 4 & 079-60606600/ 1/ 2
e-mail: abad@resonance.ac.in

PATNA

Tel.: 0612-3192222, 3192222/ 3
e-mail: patna@resonance.ac.in

JODHPUR

Tel.: 0291- 6060660
e-mail: jodhpur@resonance.ac.in

AJMER

Tel.: 0145-3192222, 6060660/ 65
e-mail: ajmer@resonance.ac.in

INDORE

Tel.: 0731-3192222, 4274200
e-mail: indore@resonance.ac.in

SIKAR

Tel.: 01572-319222, 606066
e-mail: sikar@resonance.ac.in

AGRA

Tel.: 0562-3192222, 6060660
e-mail: agra@resonance.ac.in

RANCHI

Tel.: 0651-6060660
e-mail: ranchi@resonance.ac.in

ALLAHABAD.

Tel.: 0532-6060660
e-mail: allahabad@resonance.ac.in

NASHIK

Tel.: 0253-6090028
e-mail: nashik@resonance.ac.in

RAIPUR

Tel.: 0771- 6060660
e-mail: raipur@resonance.ac.in

AURANGABAD

Tel.: 0240-6060660
e-mail: aurangabad@resonance.ac.in

JABALPUR

Tel.: 0761- 6060660
e-mail: jabalpur@resonance.ac.in

GWALIOR

Tel.: 0751-6060660
e-mail: gwalior@resonance.ac.in

CHANDRAPUR

Tel.: 07172-606066
e-mail: chandrapur@resonance.ac.in

SURAT

Tel.: 0261-6060660
e-mail: surat@resonance.ac.in

RAJKOT

Tel.: 0281-6002011
e-mail: rajkot@resonance.ac.in

VADODARA

Tel.: 0265-6060660
e-mail: vadodara@resonance.ac.in

BASE STUDY CENTRES

Base Education Service Pvt. Ltd.

Bengaluru (Main Branch):

Reg. Office : No.27, Next to Indian Oil
Petrol Bunk, Bull Temple Road,
Basavanagudi, Bengaluru- 560004
Tel. No.: 42604600/ 95381 41504
E-Mail : info@base-edu.in
Website: www.base-edu.in

BANASANKARI II STAGE

Tel.: 26710835/26710836

BELAGAVI

Tel.: 0831-4208687 | Mobile: 9845228000

CHITRADURGA

Mobile: 9886464755, 9972413844

HUBLI

Tel.: 0836-2252685 | Mobile: 9844118615

INDIRANAGAR

Tel.: 41179342/25201306

KALYAN NAGAR

Tel.: 080-25443363/25443364

KORAMANGALA

Tel.: 40925512/40925534

MALLESHWARAM

Tel.: 41400008

MYSURU

Tel.: 0821-4242100 /4258100/4243100

RAJAJINAGAR

Tel.: 08023327588/41162135

SHIVAMOGGA

Tel.: 08182-223980, 8884849590

TUMAKURU

Tel.: 0816-2252387

UDUPI

Tel.: 0820-2522449, 2522994, 9986663074

VIJAYANAGAR

Tel.: 23111333/23111334

YELAHANKA

Tel.: 08028463922/42289643

CHIKKAMAGALURU

Mobile: 7411329369, 9448396890

HASSAN

Mobile: 9481392014, 9972038283

J P NAGAR

Tel.: 26595151/26595153.

KALABURGI

Tel.: 08472-230914
Mobile: 9845905200/9844510914



Resonance[®]
Educating for better tomorrow

Corporate Office: CG Tower, A-46 & 52, IPIA, Near City Mall, Jhalawar Road, Kota (Rajasthan)- 324005

Reg. Office: J-2, Jawahar Nagar Main Road, Kota (Raj.)- 05 | Tel. No.: 0744-3192222, 3012222, 6635555 | CIN: U80302RJ2007PLC024029

To Know more: sms RESO at 56677 | E-mail: contact@resonance.ac.in | Website: www.resonance.ac.in